

## THE 3 HOURS-HYDROTHERMAL SYNTHESIS OF HIGH SURFACE AREA SUPERPARAMAGNETIC Fe<sub>3</sub>O<sub>4</sub> CORE-SHELL NANOPARTICLES

Esty Octiana Sari, Ahmad Fadli and Amun Amri

Department of Chemical Engineering Riau University  
Jl. HR Subrantas KM 12,5 Panam, Pekanbaru 28293, Riau  
E-mail: [octiana89@gmail.com](mailto:octiana89@gmail.com)

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### ABSTRACT

**THE 3 HOURS-HYDROTHERMAL SYNTHESIS OF HIGH SURFACE AREA SUPERPARAMAGNETIC Fe<sub>3</sub>O<sub>4</sub> CORE-SHELL NANOPARTICLES.** The monodisperse core-shell Fe<sub>3</sub>O<sub>4</sub> nanoparticles have been successfully synthesized by short times (3 hours) hydrothermal method at 220 °C from FeCl<sub>3</sub>, citrate, urea and PEG. The as-synthesized samples have been characterized using X-Ray Diffraction (XRD), Transmission Electron Microscope (TEM), Brunauer-Emmet-Teller (BET) surface area analyzer, and Vibrating Sample Magnetometer (VSM). The XRD result shows the as-synthesized products are pure Fe<sub>3</sub>O<sub>4</sub>. The TEM image shows the magnetite nanoparticles have monodisperse core-shell shape. The BET result shows the magnetite nanoparticles have 650.757 m<sup>2</sup>/g surface area. The hysteresis curve shows the magnetite nanoparticles exhibit super paramagnetic properties. This simple method obtained 60 nm core-shell Fe<sub>3</sub>O<sub>4</sub> particles with super paramagnetic, high surface area as well as hydrophilic properties. Those properties are promising for various biomedical application. The advantages of simple and short times methods with high quality of product make this method very promising to be applied.

**Keywords:** Core-shell, Hydrothermal method, Superparamagnetic, Nanoparticles

### ABSTRAK

**3 JAM-SINTESIS HIDROTHERMAL NANOPARTIKEL CORE-SHELL SUPERPARAMAGNETIK Fe<sub>3</sub>O<sub>4</sub> DENGAN LUAS PERMUKAAN TINGGI.** Nanopartikel monodispersi Fe<sub>3</sub>O<sub>4</sub> berbentuk *core-shell* telah berhasil disintesis dengan metode hidrotermal waktu singkat (3 jam) pada suhu 220 °C dari FeCl<sub>3</sub>, sitrat, urea dan PEG. Sampel hasil sintesis telah dikarakterisasi dengan menggunakan difraksi sinar-X (XRD), *Transmission Electron Microscopy (TEM)*, analisis luas permukaan *Brunauer-Emmet-Teller (BET)* dan *Vibrating Sample Magnetometer (VSM)*. Hasil XRD menunjukkan bahwa produk adalah Fe<sub>3</sub>O<sub>4</sub> murni. Citra TEM menunjukkan nanopartikel magnetit memiliki bentuk *core-shell*. Hasil BET menunjukkan nanopartikel magnetit memiliki luas permukaan 650,757 m<sup>2</sup>/g. Kurva histeresis menunjukkan partikel nano magnetit menunjukkan sifat superparamagnetik. Metode sederhana ini menghasilkan Fe<sub>3</sub>O<sub>4</sub> berbentuk *core-shell* berukuran 60 nm bersifat superparamagnetik, luas permukaan tinggi serta bersifat hidrofilik. Sifat-sifat tersebut menjanjikan berbagai aplikasi biomedis. Kelebihan metode yang sederhana dan cepat dengan kualitas produk yang tinggi membuat metode ini sangat menjanjikan untuk diterapkan.

**Kata kunci:** Core-shell, Metode hidrotermal, Superparamagnetik, Nanopartikel

### INTRODUCTION

In recent years, magnetite (Fe<sub>3</sub>O<sub>4</sub>) nanoparticles receive much attention from multidisciplinary scientists due to its potential application in various fields. Magnetite nanoparticles have been used as

ferrofluids, magnetic recording media, heavy metal removal and various medical application such as contrast agent MRI, hyperthermia and drug delivery [1-5].

The use of magnetite nanoparticles in biomedical application especially for cancer treatment has more challenge because its characteristics have to meet some criterions such as biocompatible, superparamagnetic, monodisperse, 5.5 - 200 nm in diameter [6], and high surface area. Thus, the controllable preparations of this material are needed.

Over the past decades, researchers have proposed several methods for the synthesis of  $\text{Fe}_3\text{O}_4$  nanoparticles, such as microemulsions, co-precipitation, sol-gel method, pyrolysis, hydrothermal method etc. Among them, the hydrothermal technique can provide highly crystalline  $\text{Fe}_3\text{O}_4$  nanoparticles with good magnetic properties and highly monodispersity [7, 8]. Cheng *et al.* [9] reported a facile hydrothermal method to prepare the monodisperse and high water solubility of  $\text{Fe}_3\text{O}_4$  from  $\text{FeCl}_3$ , citrate, polyacrylamide and urea at 200 °C for 12 hours. However, the agglomeration still observed and it makes the nanoparticles size is big (210 nm). This method also time consuming because it takes 12 hours to obtain products.

In this work, we use a high temperature (220 °C) and a short time (3 hours) hydrothermal process to synthesis the core-shell magnetite nanoparticles from  $\text{FeCl}_3$ , citrate and urea. We also added polyethylene glycol as capping agent to overcome agglomeration and enhance the biocompatibility [10-12].

Our method is simple, cheap and environmentally friendly approach for the controllable synthesis of  $\text{Fe}_3\text{O}_4$  nanoparticles. To the best our knowledge, the magnetite nanoparticles obtained in this work are higher in surface area rather than the magnetite nanoparticles ever prepared before [17-20]. This method also produce magnetite nanoparticles which are ideal in size for biomedical application.

## EXPERIMENTAL METHOD

### Materials and Tools

All of the chemical reagents used in this experiment are analytical grade and use without further purification. The chemical reagents for prepare magnetite are  $\text{FeCl}_3$ , urea, polyethylene glycol (PEG) ( $M_w = 950-1050$ ) 99%, sodium citrate ( $\text{C}_6\text{H}_5\text{O}_7\text{Na}_3 \cdot 2\text{H}_2\text{O}$ ), aquades and ethanol 98%. All reagents were purchased from Merck Chemical Industry.

Magnetite nanoparticles were prepared using glass utensils, teflon lined autoclave and oven. X-ray diffractometer (XRD) X'PERT POWDER PW 30/40 was used for nanoparticles identification. The morphology of nanoparticle was analyzed by TEM JEOL JEM 1400. The magnetic properties of nanoparticles were determined by *vibrating sample magnetometer* (VSM) type Oxford VSM1.2H. Surface area was analyzed using the Brunauer–Emmett–Teller (BET) tests performed with Quantachrome surface area analyzer.

## Synthesis

Synthesis of magnetite nanoparticle has been carried out using the hydrothermal method. A total of 2 mmol  $\text{FeCl}_3$  (0.05M), 4 mmol of sodium citrate (0.10 M), and 6 mmol of urea (0.15M) was dissolved in 40 mL of distilled water, then polyethylene glycol (PEG) (7.5 g/L) was added until completely dissolved. The solution was then transferred into a teflon lined autoclave which subsequently was putted into oven with temperature was set to 220 °C for 3 hours. Black precipitate was separated by a bar magnet, then washed with water and ethanol, and dried at 60 °C over the night.

## Characterization

The crystal structure, purity and crystal size of  $\text{Fe}_3\text{O}_4$  were determined by XRD. Approximately 200 mg of sample was printed directly on  $2 \times 2.5$  cm of aluminum. Samples were characterized using XRD instrument with Cu radiation light ( $\lambda = 1.54 \text{ \AA}$ ) with a voltage of 40 kV and current of 30 mA at  $2\theta = 10-99.97^\circ$ . The particle size, morphology and aggregate were analyzed by transmission electron microscope (TEM JEOL JEM 1400) operated at 5 kV under vacuum. Surface area was analyzed using the Brunauer-Emmett-Teller (BET) test performed with Quantachrome surface area analyzer. The magnetic characteristic of magnetite nanoparticles was analyzed using vibrating sample magnetometer (VSM) type VSM1.2H Oxford. Magnetization curves were measured at room temperature with a magnetic field that varies in the range -1 to 1 Tesla.

## RESULTS AND DISCUSSION

The crystallographic structure of the sample was identified by XRD. The XRD pattern of the typical sample is depicted in Figure 1. It is found that the peaks in XRD pattern match well with the  $\text{Fe}_3\text{O}_4$  (ICSD no.01-089-0950). However, there are not typical  $\epsilon\text{-Fe}_2\text{O}_3$  peaks such as (013), (122), (210) (ICSD no.01-076-8881) existing in the XRD pattern, confirm the as-synthesized sample are pure  $\text{Fe}_3\text{O}_4$  phase.  $\epsilon\text{-Fe}_2\text{O}_3$  is a ferromagnetic orthorhombic iron oxide, while  $\text{Fe}_3\text{O}_4$  nanoparticle is superparamagnetic cubic iron oxide [21]. The superparamagnetic properties are expected for drug delivery application to overcome particles agglomeration because of residual magnetization after exposed by external magnet [22].

The calculation using Scherrer's formula for the strongest peak (3 1 1) reveals that the crystal size of the sample is 9.68 nm. The Miller indices ((311) and (440)) are combination of all odds or all even numbers, indicating the crystal structure of as-synthesized sample is Face Centered Cubic (FCC) [16].

The morphologies and particle size of typical sample were examined by TEM. The TEM images in

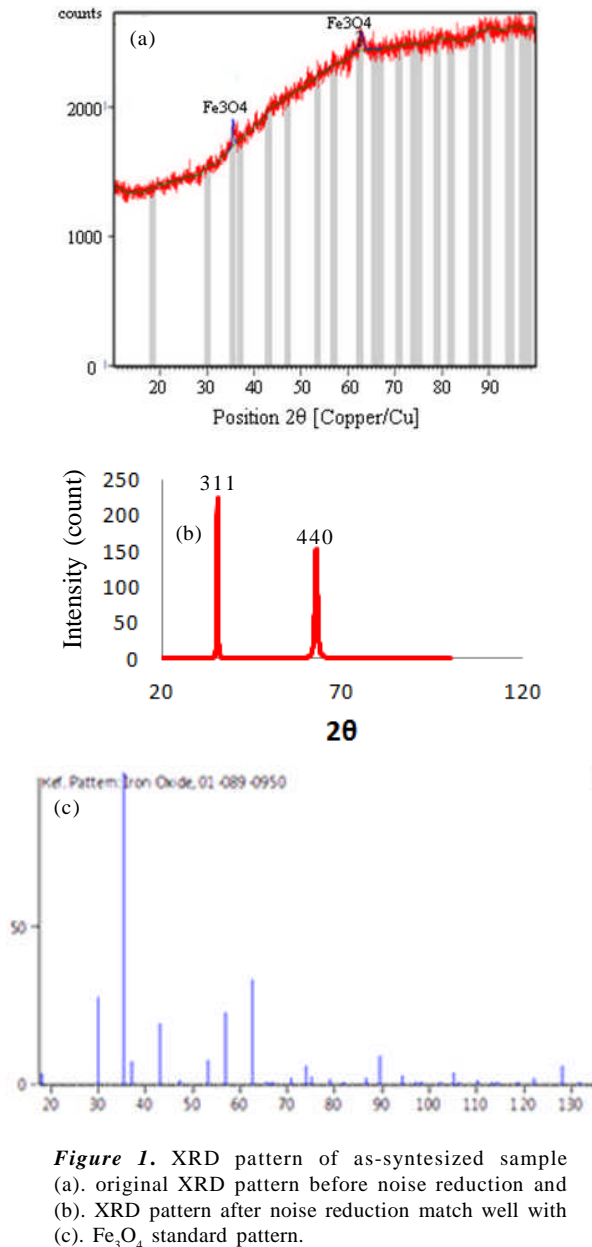


Figure 2(a) and 2(b) show that the product are monodispers and no agglomeration detected. In high magnification (Figure 2c and 2d) it can be clearly observed that the particles have a core-shell shape. It is identified from the dark image in core and the light image in shell area. The dark image indicates a higher christallinity, meanwhile the light image indicates an amorphous phase [9]. The core-shell shape gives some advantages, the shell can be act as protector from air oxidation and can be place for ligand attachment as well [13-14]. From the TEM images, the average particle size of the as-synthesized sample is about 60 nm. This size is suitable to the criteria of ideal drug delivery where the particles size must be in range of 5 - 200 nm. It is reported that under 5 nm the drug delivery will be released fastly by renal, and if it is bigger than 200 nm it will be deposited in liver [6].

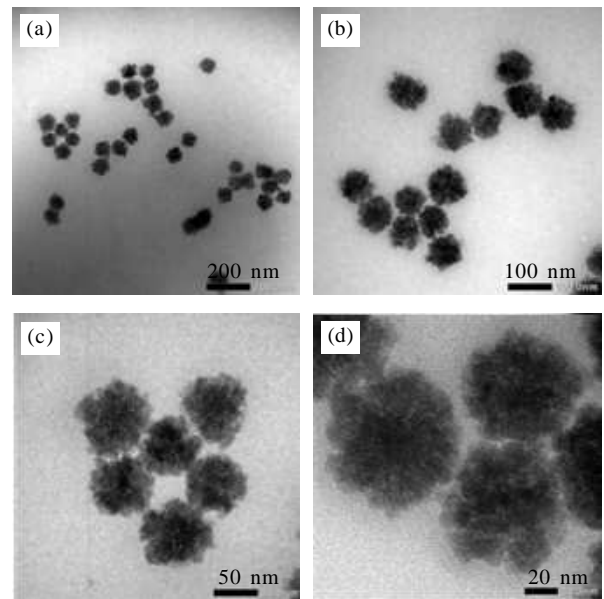


Figure 2. low magnification (a and b) and high magnification (c and d) TEM image of typical sample.

Surface area of sample was determined by the Brunauer-Emmet-Teller (BET) test. The BET plot is shown in Figure 3. From the calculation, the specific surface area of sample is  $650.575 \text{ m}^2/\text{g}$ . This surface area is greater than the surface area of magnetite nanoparticles from literatures. It was reported that Cheng et al. have prepared magnetite with the specific surface area was  $56.7 \text{ m}^2/\text{g}$  [9]. Rahman et al. have prepared magnetite with the specific area was  $23.1 \text{ m}^2/\text{g}$  [17]. Mascolo et al. prepared magnetite with the specific surface area were  $89.4\text{-}167.9 \text{ m}^2/\text{g}$  [18]. Ma et al. prepared magnetite with the specific area was  $286.9 \text{ m}^2/\text{g}$  [19]. On 2016, Jianfang et al. prepared magnetite with the specific area was  $68 \text{ m}^2/\text{g}$  [20].

The magnetic properties of as-synthesized sample were measured on Vibrating Sample Magnetometer (VSM) at 300K. The magnetic hysteresis

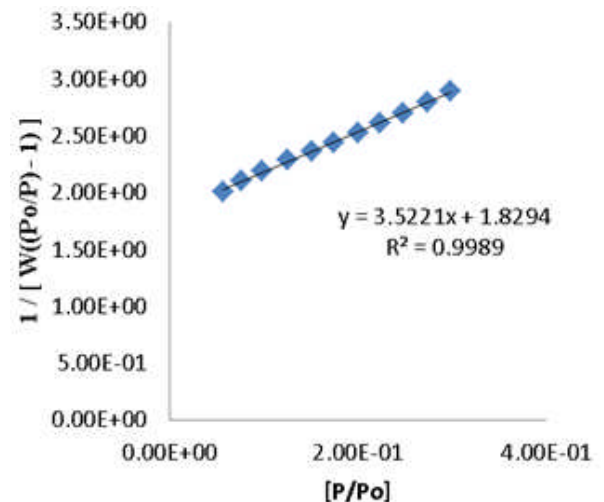


Figure 3. BET curve of typical sample

curve depicted in Figure 4 shows no coercivity and slightly hysteresis loop indicate a typical superparamagnetic sample. This superparamagnetic behavior gives some advantages in biomedical application, such as easy to be controlled and guided to the parts of body [15].

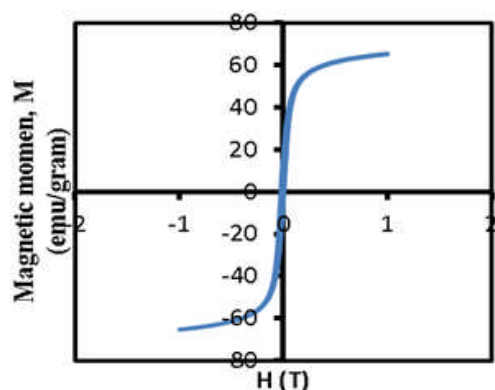


Figure 4. Magnetic hysteresis curve of typical sample

Not only has superparamagnetic property, the as-synthesized magnetite is also water soluble (Figure 5). The presence of PEG that has an -OH group on the surface of the nanoparticles makes the nanoparticles are hydrophilic and negatively charged. Thus, it makes the nanoparticles have a longer residence time in the bloodstream (high bioavailability) because it corresponds to the nature of the plasma proteins which are also hydrophilic and negatively charged. In general, phagocytosis will occur in particles of non-polar and positively charged as it is considered as a substance foreign to the body [6].

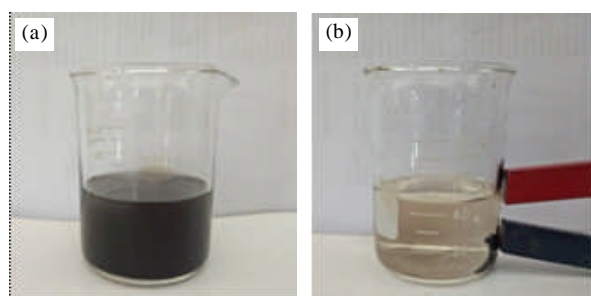


Figure 5. Photograph of the as-synthesized core-shell magnetite nanoparticles dispersed in water (a) without magnet and (b) after applied an external magnetic field for 45 seconds explain.

## CONCLUSION

The magnetite core-shell nanoparticles have been successfully synthesized from environmental friendly starting material  $\text{FeCl}_3$ , citrate, urea and polyethylene glycol (PEG) using a facile and economical hydrothermal method. The core-shell magnetite synthesized in this process has ideal size for drug delivery (60 nm), high surface area ( $650.575 \text{ m}^2/\text{g}$ ), water soluble and exhibit

superparamagnetic properties. Those all properties will render them as ideal candidates for biomedical applications.

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